

Vibrational Studies of Calcium Magnesium Ultraphosphate Glasses

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ABSTRACT

Phosphate glasses in a system of magnesium calcium phosphate for host matrix luminescent materials were investigated. The glass samples in a series of $x\text{MgO}-(40-x)\text{CaO}-60\text{P}_2\text{O}_5$ in molar ratio ($0 \leq x \leq 40$ mol%) were prepared using melt-quenching technique. The effect of Mg and Ca content on the structure and glass stability against crystallization in the system of $\text{MgO}-\text{CaO}-\text{P}_2\text{O}_5$ ultraphosphate has been investigated using XRD, Infrared and Raman spectroscopy. The compositions containing up to 40 mol% of CaO and/or MgO formed glasses and no crystalline phases were detected by XRD. The results of IR and Raman spectroscopy suggest that the phosphate network of these glasses is composed of middle (Q^2) and branching (Q^3) phosphate tetrahedral and other calcium/magnesium anions. All the symmetric and asymmetric stretching vibrations of PO_4 and PO_2 observed in the spectra are characteristic of Q^3 and Q^2 groups. The glass network, especially the Q^2 units can be modified by the presence of Ca and Mg ions. This modification is primarily associated with the phosphate and the Q^2 and Q^3 phosphate units randomly distributed in the network. Spectroscopic results shows that the modification of the phosphate network is higher for the Ca containing glasses with respect to the Mg ones, at the same alkali earth content, due to the well defined Ca properties as a modifying cation. Formation of $\text{P}-\text{O}-\text{H}$ bond expresses the hygroscopic nature of the phosphate glasses. This study shows that the vibrational spectroscopy (Infrared and Raman) are provide useful, complementary information about the network structures of ultraphosphate glasses.

| X-ray diffraction | Infrared spectroscopy | Raman spectroscopy | Phosphate glasses | Ultraphosphate |

1. Introduction

Glasses are increasingly used as host material for solid state lasers based on rare earth and transition metal ionic transitions. The advantages with glasses are the ease with which large homogeneous pieces can be produced. The nonlinear refractive indices can be made low and the doping concentration can be large. It has been shown that the fluorescence intensity of rare earth ion are enhanced several fold when glass hosts with low phonon energies as example in germanate and tellurite are used [1,2].

In recent years, phosphate based glasses have been interest due to their several special properties such as large thermal expansion coefficients, low melting temperatures, solubility, etc. [3,4]. It is also an excellent host material due to good chemical durability, ion exchange ability, high gain coefficient, wide band width capability and low upconversion emission. The nonradiative processes include processes characteristic of the host and processes characteristic of the active ion promoted by its concentration. Phosphate glass present high solubility of rare earth ions allowing a high concentration of active ion into a small volume of the phosphate glass [5]. Among oxide glasses, phosphate glasses are particularly attractive hosts because they can accommodate large concentrations of active ions without losing the useful properties of the material. Phosphate glasses are relatively easy to prepare and offer an important range of compositional possibilities (ultra, meta, pyro, and ortho-phosphate), which facilitate tailoring of the physical and chemical properties of interest for specific technological applications [6,7].

The basic structural units of crystalline and amorphous phosphates are the PO_4 tetrahedral and these tetrahedral links through covalent bridging oxygen to form various phosphate anions [8-11]. These tetrahedral are classified using the Q^n terminology, where n represents the number of bridging oxygen per tetrahedron [12,13]. This notation is usually use in NMR spectra. The networks of phosphate glasses can be classified by the oxygen-to-phosphorus ratio, which sets the number of tetrahedral linkages, through bridging oxygen, between neighboring P-tetrahedral. Phosphate glasses can be made with a range of structures, from a cross-linked network of Q^3 tetrahedral (vitreous $v\text{-P}_2\text{O}_5$) to polymer-like metaphosphate chains of Q^2 tetrahedral to pyro- (Q^1) and orthophosphate (Q^0) anions, depending on the modifier composition as shown schematically in Fig. 1.

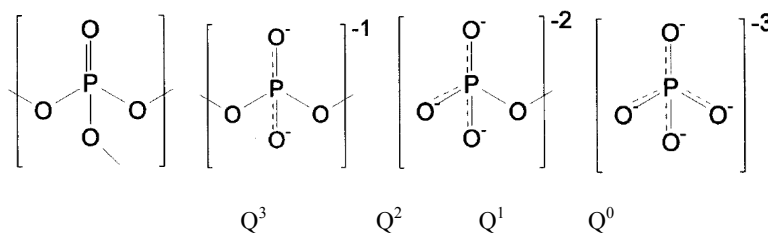


Figure 1: P-tetrahedral sites that can exist in phosphate glasses. [31]

Spectroscopy (IR, Raman and NMR) has been used as an important tool to study the nature of glasses for the past many years. The relative impact of spectroscopy has probably been greater in the analysis of glasses than those of crystals. Galeener and Mikkelsen [14] reported the first Raman spectra from anhydrous $v\text{-P}_2\text{O}_5$. Due to the P_2O_5 is hygroscopic and volatile, anhydrous glasses have been prepared using dry-box handling procedures from crystalline P_2O_5 and then melted in sealed silica ampoules. Raman spectra from P_2O_5 glass as shown in Fig. 3 are dominated by bands near 1380 cm^{-1} , due to stretching modes of bridging oxygen $\text{P}=\text{O}$, and at 790 cm^{-1} , due to symmetric $\text{P}-\text{O}-\text{P}$ stretching symmetry modes. The intense IR absorption band, around 930 to 970 cm^{-1} is assigned to asymmetric $\text{P}-\text{O}-\text{P}$ stretching modes [15]. Phosphorus-oxygen, $\text{O}=\text{PO}_3$, tetrahedral bonded randomly in a three-dimensional network where in each silicon atom is bonded with four phosphorus atoms by oxygen linkages, and each phosphorus atom has only three such bridging bonds (Fig.1). In the fourth vertex of the $\text{O}=\text{PO}_3$ tetrahedron there is a non-bridging oxygen atom bound with the central phosphorus atom with a double bond, $\text{O}=\text{P}$. Owing to a high polarizability of the double bond an intensive band with a maximum at 1380 cm^{-1} caused by the stretching vibrations of the $\text{O}=\text{P}$ bonds arises in the Raman spectra [10-13].

The addition of a modifying oxide to $v\text{-P}_2\text{O}_5$ results in the creation of non-bridging oxygen (NBO) at the expense of bridging oxygen (BO). The resulting depolymerization of the phosphate network with the addition, e.g., of alkali earth oxide, RO , can be described by the pseudo-reaction $2Q^n + \text{RO} \rightarrow 2Q^{n-1}$ [13]. For binary

$x\text{RO}-(1-x)\text{P}_2\text{O}_5$ glasses, the concentrations of Q^n tetrahedral (Fig. 5) generally follow simple predictions based on the glass composition. In ultraphosphate region, the fractions of Q^2 and Q^3 tetrahedra are given by NMR study as $f(Q^2) = x/(1-x)$ and $f(Q^3) = (1-2x)/(1-x)$ [13]. Based on these equation, the composition of this study ($x=0.40$), the ratio of Q^3 to Q^2 is 2/3. The evolution of the cross-linked Q^3 network of P_2O_5 to the chain-like Q^2 network of a metaphosphate glass has been seen clearly in the ^{31}P NMR spectra [16,17]. The bridging-to-terminal oxygen ratio, for $x\text{RO}-(1-x)\text{P}_2\text{O}_5$ depends on composition according to $BO/TO = 0.5(3-4x)$ [16].

The understanding of the glasses structure on different length scales remains difficult mainly due to the lack of periodicity inherent to glasses. Only the suitable correlation of experimental data obtained from several investigations provided by different methods allows the finding of helpful structural information. The use of complementary techniques such as Raman, infrared (IR) and NMR spectroscopy has proven to be an efficient examination way of the glasses structure, in general and of the phosphate glasses structure, in particular. In this study, basic composition of 60% P_2O_5 , 0-40% CaO and 0-40% MgO (all the compositions are referred to mol %), are chosen as host base glass. The aim of the present work is to investigate the complementary information about the local order around phosphorus tetrahedral structures due to the incorporation of mix alkali earth cation (Ca and Mg) oxide to ultraphosphates network using both Fourier-transformed infrared (FT-IR) and Raman spectroscopy. The term ultraphosphate refers to compositions where $O/P < 3.0$. Ultraphosphate glasses in this study, $x\text{RO}-(1-x)\text{P}_2\text{O}_5$ systems have $x < 0.5$.

2. Experimental

Sample Preparation

A nominal composition of the glass sample $x\text{MgO}-(40-x)\text{CaO}-60\text{P}_2\text{O}_5$ in mol% (with $0 \leq x \leq 40$), were prepared using melt-quenching technique. The batch mixture (20g) was prepared using raw materials of reagent grade CaCO_3 , MgO and H_3PO_4 (85% liquid). The corresponding weights of the starting materials were mixed thoroughly in platinum crucible and placed in air in an electric furnace. A three-step melting schedule was employed to minimize volatilization losses of low-melting starting materials as well as to ensure homogeneous mixing of the constituents. At the first step, the temperature was set to 100 °C with heating rate of 5 °C/min, and maintained for 1 h to facilitate evaporation of water released by P_2O_5 . The temperature was then raised to 300 °C (heating rate = 20 °C/min) and left for another 1 h to release CO_2 . In the third step, the temperature was increased slowly from 900 to 1250 °C with heating rate of 10 °C/min and kept there for 2 h in air for the glass to melt. The melts were quickly cooled at room temperature by pouring and pressing between two stainless steel plates to obtain the glass.

Sample Characterization

The structures of the crystalline and amorphous prepared samples were analyzed by means of X-ray diffraction measurements (XRD), using powders form. The XRD measurements were carried out with CuK_α radiation operating at 40 kV, 30 mA with Bragg-Brentano geometry at room temperature using Siemens Diffractometer D5000, equipped with diffraction software analysis. Diffraction patterns were collected in the 2-theta (2θ) range from 20 to 80°, in steps of 0.04° and 4s counting time per step.

To obtain IR absorption spectra of a glass sample, KBr pellet technique is employed. The glass samples were ground in a clean mortar to a fine powder and weighed quantity (~4 mg) of the powder was intimately mixed with desiccated highly purified KBr powder. The mixture was then pressed with a pressure of 5 ton per square

inch to yield a transparent pellet of approximate thickness 0.1 mm suitable for mounting in the spectrophotometer. All infrared spectra of the glasses were recorded using a Perkin-Elmer Spectrum One FT-IR spectrometer over the range of wavenumber $4000\text{--}400\text{ cm}^{-1}$ at room temperature using 100 scans at 4 cm^{-1} resolution.

The Raman spectra were measured using Perkin-Elmer model 2000R NIR FT-Raman spectrometer in the spectral range $100\text{--}4000\text{ cm}^{-1}$, with the laser power on the samples being 250 mW. The spectrometer uses an (Nd:YAG) as a lasing medium. The laser is software-controlled, diode-pumped running at 810 nm, with power of 1600 mW or 750 mW TEM00 and typical working power range from 100 mW to 1 W. The resultant Gaussian beam profile is vertically polarized to better than 100:1. The amplitude stability is better than 0.1 rms and is controlled by means of active optical feedback. The laser line width gives a spectral resolution of better than 0.5 cm^{-1} .

3. Results

The calcium magnesium phosphate glass samples obtained were transparent and showed no signs of devitrification. XRD spectra of the studied glasses have shown that all the samples are completely amorphous and they do not present any crystalline phase (Fig. 2).

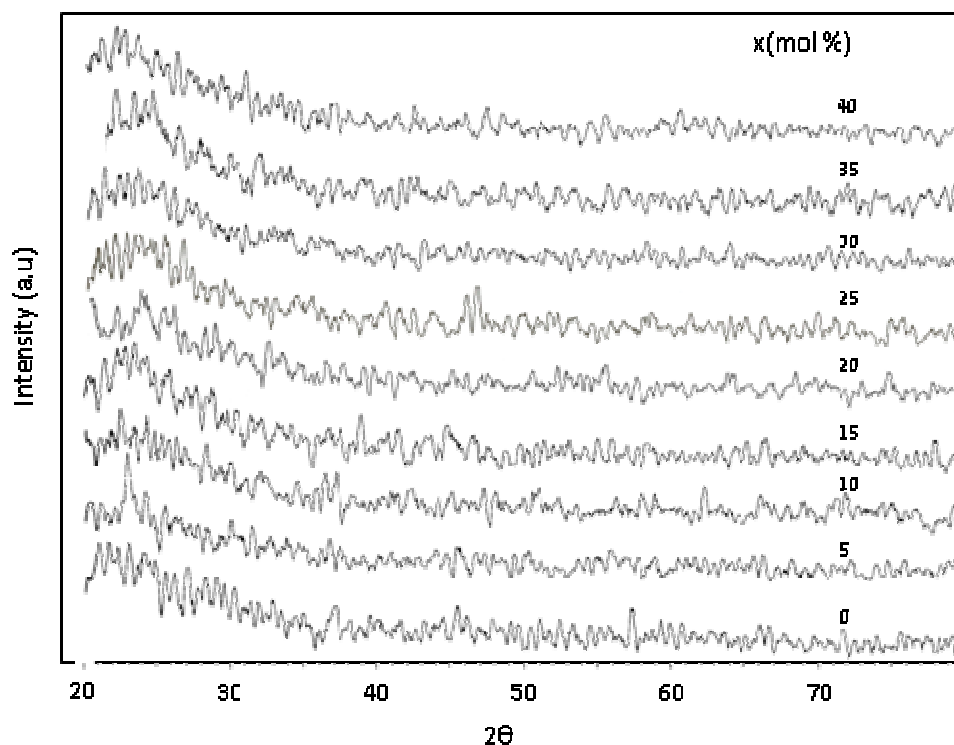


Figure 2: XRD patterns of $x\text{MgO}-(40-x)\text{CaO}-60\text{P}_2\text{O}_5$ glasses, with $0 \leq x \leq 0.4$.

3.1 Raman Spectra

Raman spectra of $x\text{MgO}-(40-x)\text{CaO}-60\text{P}_2\text{O}_5$ with compositions in the range $0 \leq x \leq 40$ mol% are shown in Fig. 3. The most important bands and their assignments are given in Table 1. The assignment of the peaks was done following guidelines of alkali (Li, Na) and (Cu, Zn) ultraphosphate glasses [8,9]. The shapes of Raman spectra of the investigated glasses (Fig. 3) are similar and feature a dominant peaks at around 1320 cm^{-1} and 1180 cm^{-1} , a shoulder with low intensity around $\sim 1100\text{ cm}^{-1}$, a shoulder at 792 cm^{-1} , peaks at $\sim 662\text{ cm}^{-1}$ and a number of bands with weak intensity at 550 cm^{-1} , 455 cm^{-1} and $\sim 300\text{ cm}^{-1}$. The peaks around 1320 cm^{-1} is due to the P=O symmetric stretch. Bands near 1180 cm^{-1} are due to the asymmetric stretch of PO_2 groups, i.e. stretching motions of the two nonbridging oxygen (NBO) atoms bonded to phosphorous atoms (PO_2) in the Q^2 phosphate tetrahedral (in the Q^n terminology, n represents the number of bridging oxygen (BO) per PO_4 tetrahedron (will discussion below), and those near 1080 cm^{-1} are due to the symmetric stretch of PO_2 groups. Bands in the frequency region from 655 to 670 cm^{-1} have been attributed to motions of bridging oxygen in P–O–P chains, originates from the P–O–P symmetric stretch of bridging oxygen in Q^2 species whereas the band at 762 – 792 cm^{-1} arises from the asymmetric vibrations of the same species. The broad band at $\sim 300\text{ cm}^{-1}$ has been assigned to bending vibrations of phosphate network/polyhedral, $\delta(\text{POP})$. The relative intensity of the P=O band (1320 cm^{-1}), the intensity of both asymmetric (1180 cm^{-1}) and symmetric (1080 cm^{-1}) PO_2 bands slightly decreases with increasing CaO content. All the frequency of P=O, symmetric and asymmetric PO_2 , and symmetric and asymmetric P–O–P band decreases with the addition of CaO.

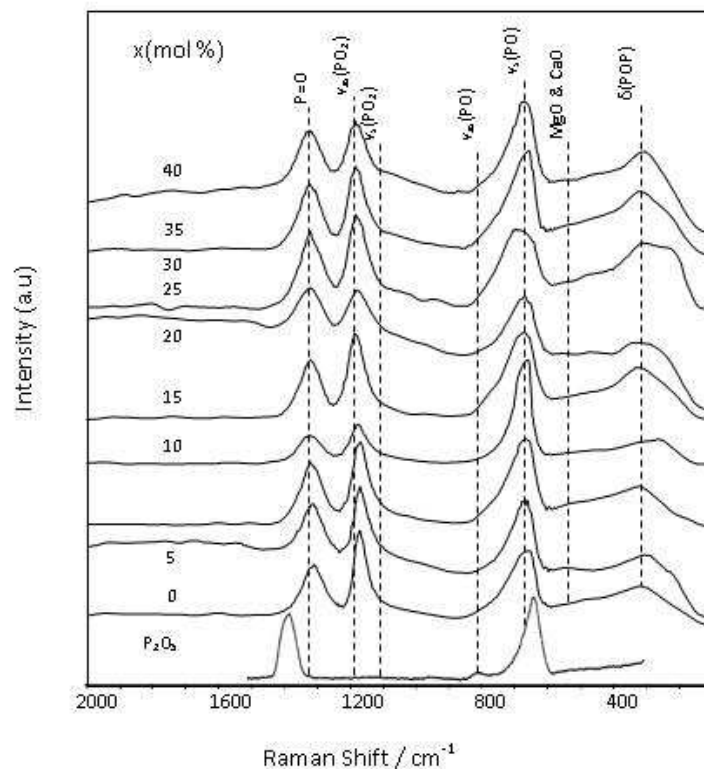


Figure 3: Raman spectra of $x\text{MgO}-(40-x)\text{CaO}-60\text{P}_2\text{O}_5$ with $0 \leq x \leq 0.4$.

Table 1: Raman bands and band assignment for $x\text{MgO}-(40-x)\text{CaO}-60\text{P}_2\text{O}_5$ with $0 \leq x \leq 0.4$.

Frequencies observed and assignments according to Refs [18,19]															
IR								Raman							
x	$\delta(\text{MgO} \& \text{CaO})$	$\nu_s\text{POP}$	$\nu_{as}\text{POP}$	νPO	$\nu_s\text{PO}_2$	$\nu_{as}\text{PO}_2$	$(\text{P=O})_s$	$\delta(\text{POP})$	$\delta(\text{MgO} \& \text{CaO})$	$\nu_s\text{POP}$	$\nu_{as}\text{POP}$	νPO	$\nu_s\text{PO}_2$	$\nu_{as}\text{PO}_2$	$(\text{P=O})_s$
P_2O_5	-	-	-							790	960				1380
40	500	640	746	973	1076	1180	1331	302	453-545	670	789	970	1089	1184	1325
35	498	643	749	950	1067	1175	1315	314	414-533	654	768	954	1086	1182	1323
30	520	650	730	970	1065	1160	1320	306	404-470	671	780	965	1085	1181	1323
25	509	655	745	967	1079	1156	1281	284	466-551	663	774	969	1082	1177	1318
20	551	662	749	937	1068	1176	1316	322	455-550	662	762	963	1080	1176	1316
15	524	665	727	963	1075	1157	1278	314	429-588	664	775	970	1079	1171	1314
10	511	660	752	951	1062	1163	1331	316	417-511	664	780	969	1077	1169	1312
5	510	620	755	967	1063	1175	1319	300	361-543	660	785	963	1074	1167	1311
0	509	659	758	956	1088	1124	1233	315	418-483	655	792	953	1073	1172	1310

3.2 Infrared Spectra

The infrared spectra of the $x\text{MgO}-(40-x)\text{CaO}-60\text{P}_2\text{O}_5$ glasses are shown in Fig. 4. The most important bands and their assignments are also listed in Table 1. The typical features of the IR spectrum of the $\nu\text{-P}_2\text{O}_5$ glass are $\nu(\text{P}=\text{O})$ band at 1380 cm^{-1} , the asymmetric stretching band, $\nu_{\text{asym}}(\text{POP})$ at 960 cm^{-1} and symmetric stretching bands, $\nu_{\text{sym}}(\text{POP})$ at 790 cm^{-1} and no intensity between 1350 and 1200 cm^{-1} as reported by Mayer [15] are also shown for comparison in Fig. 4. The IR spectra of glasses are consisting of two main sets of infrared absorption bands in the low- and high-frequency regions. The broad bands appear around 3400 and 3150 cm^{-1} . These bands can be accounted for by a different role of H_2O molecules in the structure. In particular, the atmospheric moisture is easily absorbed by the phosphate sample or by the pellet, causing the appearance of infrared band belonging to H_2O molecules although the sample under investigation does not contain H_2O as unit in the network [18]. The three bands located in the range $2700\text{--}2950\text{ cm}^{-1}$ are relatively weak and can be ascribed to the stretching vibrations of $\text{P}\text{--}\text{O}\text{--}\text{H}$ group in different structural sites. This group forms the strongest hydrogen bonding with the non-bridging oxygen. The strong band at about 1600 cm^{-1} can be attributed to the $\text{O}\text{--}\text{H}$ bending vibration [18].

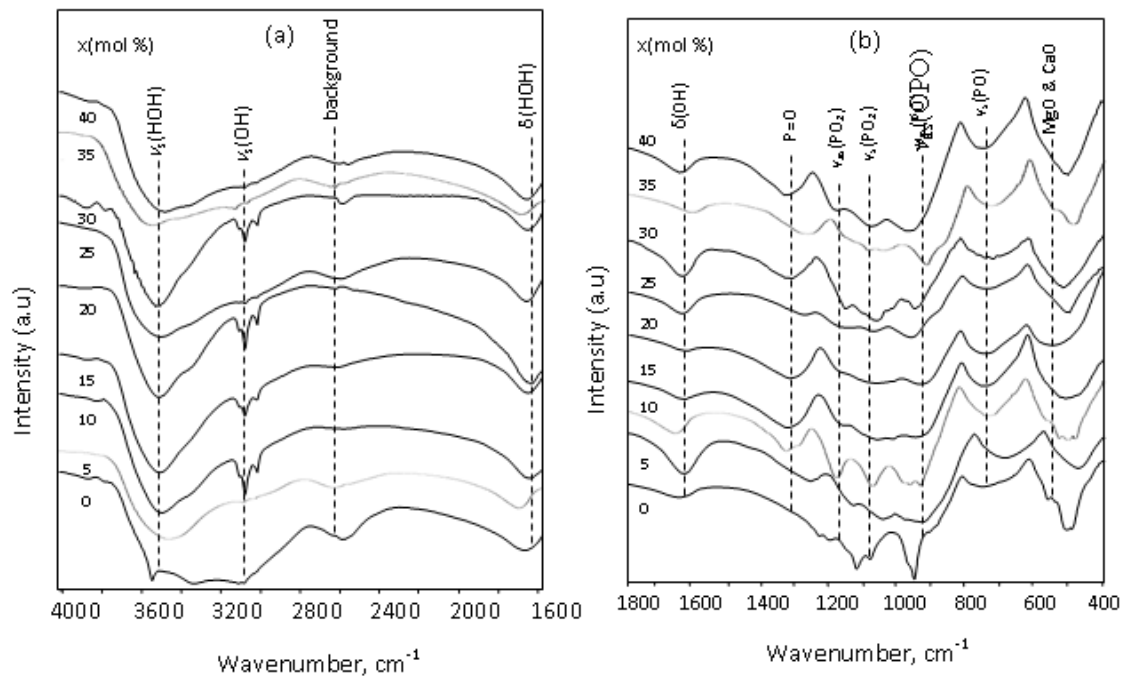


Figure 4: FT-IR spectra of the investigated glasses of $x\text{MgO}-(40-x)\text{CaO}-60\text{P}_2\text{O}_5$, with $0 \leq x \leq 0.4$ (a) IR spectra in the range of $1600\text{--}4000\text{ cm}^{-1}$ and (b) IR spectra in the range of $400\text{--}1800\text{ cm}^{-1}$.

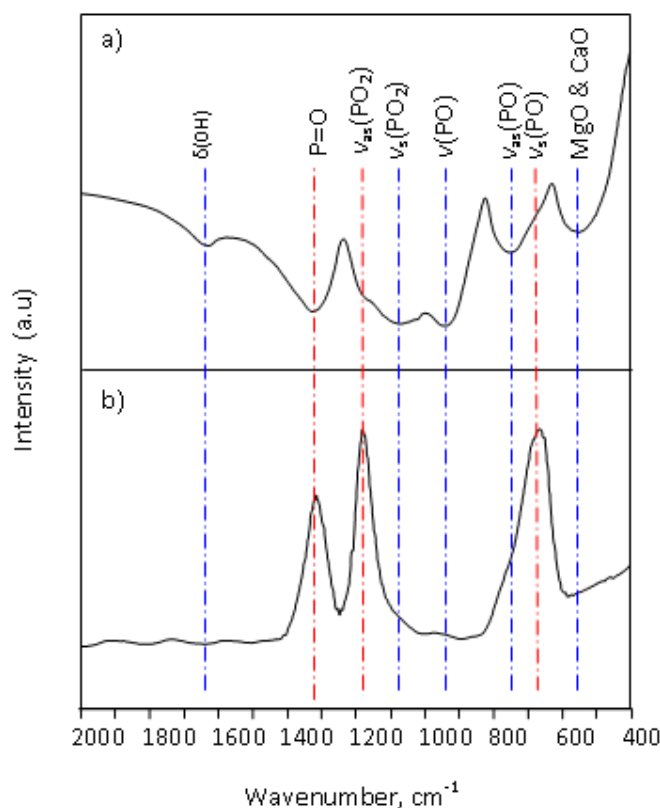


Figure 5: Vibrational spectra of 20MgO-20CaO-60P₂O₅ glass sample, (a) Infrared spectra, (b) Raman spectra.

The bands arising from the vibrations of the phosphate glass network appear in the range 1400–500 cm⁻¹ (Fig. 4(b)). Three main regions can be distinguished in this range; the region between 1400–1150 cm⁻¹ is a characteristic of vibrations of non-bridging PO₂ groups, the region around 1150–900 cm⁻¹ is characteristic of terminal P–O (and/or PO₃ groups) and the region between 900 and 700 cm⁻¹ is characteristic of the vibrations of bridging P–O–P groups. All the symmetric and asymmetric stretching vibrations observed in the spectra are characteristic of Q³ and Q² groups.

The assignments for infrared spectra as shown in Fig. 4(b) are more difficult, especially in the range 800–1200 cm⁻¹ due intensity of true peak position, location of peak shoulder and low peak intensity. In order to solve this problem, comparisons between Raman with infrared spectra have been done. Comparisons of Raman and infrared spectra for ultraphosphate glasses are shown in Fig. 5. This figure shows that the advantage and disadvantage of both technique (Raman and infrared) and their complementary each other.

The strong band observed at 1330 cm⁻¹ with its shoulder at 1370 cm⁻¹ is assigned to the characteristic stretching mode of the P=O bond. Several previous studies [19-22] have shown that the band corresponding to the stretching vibration of doubly bonded oxygen could be found in the frequency range 1230–1390 cm⁻¹. The band obviously becomes smaller (x=40), indicating a decrease in the double bond character and in the effective force constant of the (P=O) bond as found in ultraphosphates with increasing the alkali earth content. The band or

shoulder at about 1160 cm^{-1} is ascribed to the asymmetric stretching vibration of O–P–O groups, $\nu_{\text{as}}(\text{PO}_2)$, i.e., to asymmetric stretching modes of the two non-bridging oxygen atoms bonded to phosphorus atoms or Q^2 units, in the phosphate tetrahedral. Their intensity obviously increases with increasing Mg content. This result indicates that the phosphate linkages are shortened as incorporate into the glass structure and leading to increase the relative content of the Q^2 units. The medium broad band at 1070 cm^{-1} is related to the symmetric stretching vibration of that groups, $\nu_{\text{s}}(\text{PO}_2)$ [23,24]. Previous spectroscopic study [24] reported the asymmetric stretching vibration of the metaphosphate group, $\nu_{\text{as}}(\text{PO}_3)$ end groups (Q^1), in the range $1080\text{--}1180\text{ cm}^{-1}$. Therefore, in the present IR spectra, the $\nu_{\text{as}}(\text{PO}_3)$ could interfere with the band at 1160 cm^{-1} . The broadening character observed in the region $900\text{--}1100\text{ cm}^{-1}$ may lead to the assumption that the $\nu_{\text{as}}(\text{PO}_3)$ could also interfere in the range $1000\text{--}1060\text{ cm}^{-1}$ [24].

The bands/shoulder around 937 cm^{-1} is assigned to the asymmetric stretching vibration of P–O–P linkages, $\nu_{\text{as}}(\text{P–O–P})$, while the relatively weak band around 750 cm^{-1} is due to the symmetric stretching vibration of that linkages, $\nu_{\text{s}}(\text{P–O–P})$, stretching of the bridging oxygen atoms bonded to a phosphorus atom in a Q^2 phosphate tetrahedron [23,24]. The asymmetric stretching band of (P–O–P) near 920 cm^{-1} initially shifts to higher frequencies as the amount of the Mg increases. The larger wavenumber of the (P–O–P) band is a result of the smaller (P–O–P) bond angle, which results from shorter phosphate linkages or smaller metal cation size. The phosphate linkages of the glasses with higher Mg content are shorter due to the depolymerization of the glass structure.

On the basis of previous observation [22], the shoulder at about 600 cm^{-1} can be assigned to a covalent bond between non-bridging oxygen and calcium ions as P–O–Ca stretching vibration. The absorption band at about 550 cm^{-1} may be assigned either to the harmonics of P–O–P bending vibration or is assigned to harmonics of bending vibration of O=P–O linkages. This band appears in the most binary and ternary phosphate glasses [25,10]. All the symmetric and asymmetric stretching vibrations observed in the spectra are characteristic of Q^3 and Q^2 groups.

4. Discussion

The results of this studies shown that the phosphate glass structure in the metaphosphate composition, as example in Fig.3 for $20\text{MgO--}20\text{CaO--}60\text{P}_2\text{O}_5$, consists of a main network of chains composed of Q^2 tetrahedral units, where two of the oxygen atoms are bridging (P–O–P) and two are non-bridging (PO_2). When CaO and MgO is added to this group of metaphosphate chains, only little change occurred, as it is revealed from the Raman spectra for both magnesium and calcium phosphate glasses. The P=O band (1380 cm^{-1} in TeO_2 glass) shifts to lower frequency (1316 cm^{-1}), and its intensity slightly decreases with increasing CaO and/or Mg content. These results show that the changes of the frequencies in various Raman modes are consistent with delocalization.

This assessment of " π -bonding on Q^3 tetrahedral, resulting in longer P=O bonds and shorter, and thus stronger, P–BO bonds. Kreidl hypothesized that such delocalization would result in an increased network strength due to the formation of P–O– R^{2+} alkali bridges through the P=O bonds on the Q^3 species [26]. Similar arrangements are found in the structure of a magnesium ultraphosphate crystal, [8]; the magnesium ions are coordinated to 'double-bonded' oxygen on Q^3 tetrahedral as well as non-bridging oxygen on Q^2 tetrahedral. As a consequence, the addition of an alkali earth oxide leads to the formation of new alkali-oxygen bridges even as P–O–P bridges are eliminated and increase in the strength of the covalent interconnection due to the shorter P–BO bonds. There has been some controversy about the different oxygen species in phosphate glasses, particularly as it relates to the relevance of distinguishing 'double bonded' oxygen from 'nonbridging' oxygen [13]. For example, a Q^2 tetrahedron is often represented as possessing one P=O and one P–O $^-\text{R}^+$ terminal oxygen bond, the

implication being that the double bond is shorter because of the localization of the fifth valence electron from P^{5+} . This description was supported by an early X-ray photoelectron study [13] in which the O1s spectra from various phosphate glasses were decomposed into three components, representing distinct $P=O$; $P-O^-R^+$ and $P-O-P$ species. The model of phosphate tetrahedral related in this study is shown in Fig. 6.

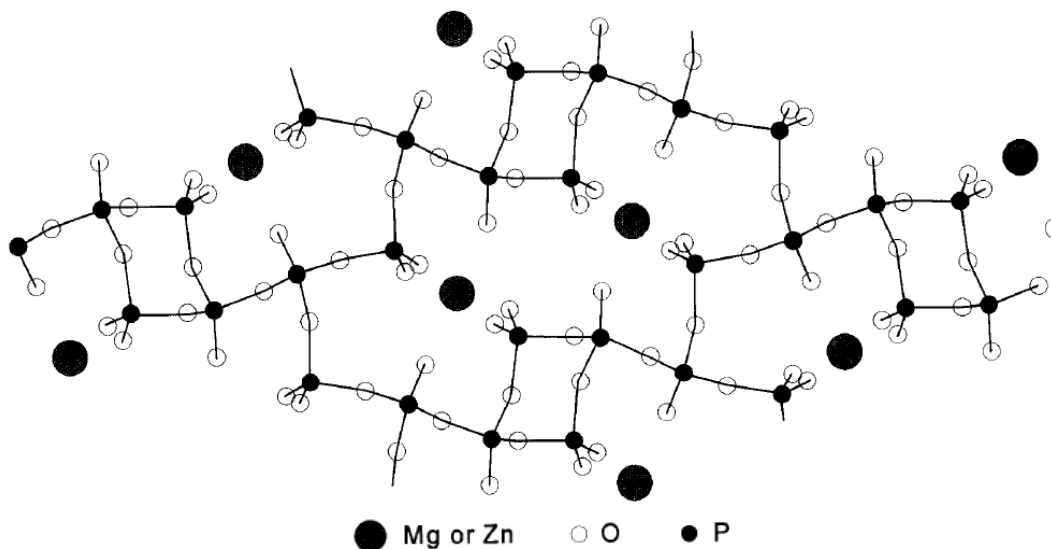


Figure 6: Example model P_2O_5 tetrahedral in the crystal structure of MgP_4O_{11} (taken from Ref. [18]).

The progression from a cross-linked Q^3 network, to chain-like Q^2 structures, with different alkali/alkali earth ratio has been monitored by a vibrational spectroscopic. As example Fig. 7 is shows the Raman spectra collected from ultraphosphate glasses with the same nominal composition of 60 mol% P_2O_5 , i.e. Li and Na-ultraphosphate [8], Zn-ultraphosphate [11] and Cu-ultraphosphate [10] glasses. The shape of Raman spectra are almost similar, shows that the structure feature in this ultraphosphate glasses are identical each other.

Bands in the $1000-1400\text{ cm}^{-1}$ range are due to the symmetric and asymmetric stretching modes of P-nonbridging (or terminal) bonds, whereas bands in the $620-820\text{ cm}^{-1}$ range are due to the symmetric and asymmetric stretching modes of bridging oxygens. For example in $(40Li_2O-60P_2O_5)$, the frequency of the dominant 'nonbridging' stretching mode decreases from 1380 cm^{-1} (in P_2O_5 glass) to 1360 cm^{-1} i.e. symmetric $P=O$ stretch on Q^3 tetrahedral, symmetric PO_2 stretch on Q^2 tetrahedral, to 1050 cm^{-1} . This decrease in frequency shows the progressive depolymerization of the phosphate network, as well as the increase in the average bond length of P-non-bridging oxygen.

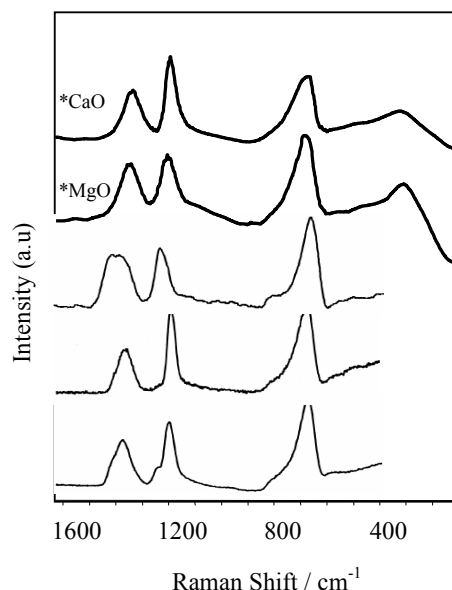


Figure 7: Raman spectra reported for ultraphosphate glasses with 40 mol% Li_2O , Na_2O [18] and ZnO [30]. These studies are shown by asterisks.

Even Raman spectra shown in Fig. 7 are similar but their vibration (peak position and width) are different in Li-, Na-, Ca-, Mg-, Cu- and Zn-ultraphosphate glasses, which are related to the expected coordination effects of the different modifiers on the phosphate network. Of particular interest is the presence of separate bands due to the $\nu(\text{P}=\text{O})$ and $\nu(\text{PO}_2)_{\text{asym}}$ modes in the spectrum from the Zn-ultraphosphate (and, to a lesser extent, the Li-ultraphosphate), whereas these two modes are indistinguishable in the spectra from the Ca- and the Na-ultraphosphate glasses. The smaller Zn^{2+} ion apparently does not form bonds to terminal oxygen $\text{P}=\text{O}$ on the remaining Q^3 tetrahedra which are as strong as those of larger Ca^{2+} and Na^+ ions, and so does not affect the frequency of the $\text{P}=\text{O}$ band to nearly the degree found for the other modifying ions.

The frequencies of the $\nu(\text{PO}_2)_{\text{sym}}$ and $\nu(\text{PO}_2)_{\text{asym}}$ peaks in Fig. 7, are cation dependent, increasing in frequency as the field strength of the modifying cation increases (in the order $\text{Na}^+ < \text{Li}^+ < \text{Ca}^{2+} < \text{Zn}^{2+}$), consistent with studies of metaphosphate glasses [27]. The full-width at half maximum (FWHM) of the $\nu(\text{PO}_2)_{\text{sym}}$ band also increases with increasing cation field strength, from about 25 cm^{-1} for the Na_2O glass to about 50 cm^{-1} for the ZnO glass, indicating an increase in the degree of disorder that may be due to changes in the distributions of the Q^2 P-O bond lengths and bond angles. Again, these observations are consistent with previous spectroscopic studies of metaphosphate glasses [28]. Phosphate glasses are composed of PO_4 tetrahedral that could be bound to one (Q^1 species), two (Q^2 species) and three (Q^3 species) other tetrahedral through the formation of P-O-P bonds at the tetrahedral corners [13]. The addition of definite amounts of metal ions, such as Ca^{2+} , Mg^{2+} , Na^+ , Li^+ , and Zn^{2+} could produce significant changes of the phosphate glass structure from a three-dimensional random network to long or very short chains of PO_4 tetrahedral.

5. Conclusions

This study shows that the vibrational spectroscopy (Infrared and Raman) provide a useful and complementary information about the network structures of ultraphosphate glasses. The structural characterization of the investigated calcium magnesium phosphate glasses by FT-IR and Raman spectroscopy revealed of the spectrum are similar to alkali/alkali earth phosphates with ultraphosphate glass composition. The IR and Raman spectra suggest that the investigated glasses have similar structures consisting of randomly distribution of Q^3 and Q^2 structural units. All the symmetric and asymmetric stretching vibrations of POP and PO_2 observed in the spectra are characteristic of Q^3 and Q^2 groups. In P_2O_5 glass, the glass structure is composed of cross-linked network of Q^3 tetrahedral with three bridging oxygen and one isolated doubly bonded oxygen, ($P=O$). With the addition of calcium oxide and/or magnesium oxide, a gradual conversion from Q^3 to Q^2 tetrahedral units occurs. The addition of definite amounts of metal ions, such as Ca^{2+} , Mg^{2+} , Na^+ , Li^+ , and Zn^{2+} could produce significant changes of the phosphate glass structure from a three-dimensional random network to long or very short chains of PO_4 tetrahedral. Spectroscopic results shows that the modification of the phosphate network is higher for the Ca containing glasses with respect to the Mg ones, at the same alkali earth content, due to the well defined Ca properties as a modifying cation.

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7. References

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